$$2.00 \times 10^5 \text{ g } \{CH_2O\} \times \frac{1 \text{ mol } \{CH_2O\}}{30 \text{ g } \{CH_2O\}} \times \frac{1 \text{ mol } O_2}{1 \text{ mol } \{CH_2O\}} = 6.67 \times 10^3 \text{ mol } O_2$$